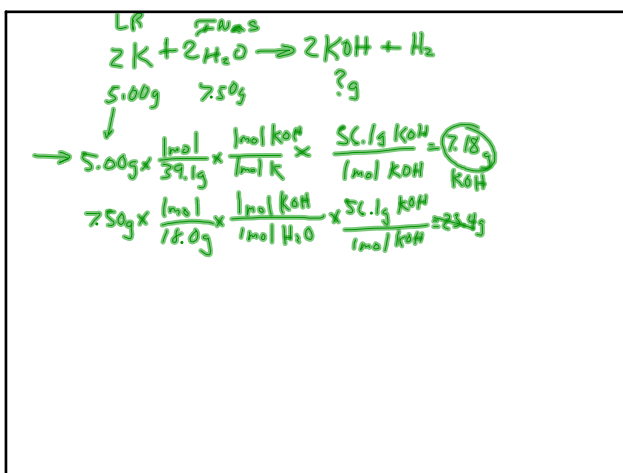


$$10.0gK \times \frac{1molK}{39.1gK} \times \frac{1molH_2O}{1molK} \times \frac{18.0gH_2O}{1molH_2O} = 4.60gH_2O$$

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$$.099gK \times \frac{1molK}{39.1gK} \times \frac{1molH_2}{2molK} \times \frac{22.4L}{1mol} = .0284L H_2$$

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→ Combined

$$\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2}$$

→ Grams

$$\frac{75.0kPa \times 25.0L}{308K} = \frac{55.0kPa \times V_2}{330K}$$

$V_2 = 36.5L$

Rate = $\sqrt{\frac{m_2}{m_1}}$

H_2 is faster $\sqrt{\frac{32}{2}}$

$O_2 = 32g/mol$
 $H_2 = 2g/mol$

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