1. $\frac{[HF]^2}{[H_2][F_2]}$

2. $\frac{[HF]^2}{(.045)(.045)} = 64$

$[HF] = .36$

3. products $K > 1$

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1. $2.0M = \frac{mol}{1.75L}$

$= 3.5 mol$

$3.5 mol \times \frac{58.5 g}{1 mol} = 205 g$

2. $M_1 \times V_1 = M_2 \times V_2$

$2.25M \times 250mL = 12.0M \times V_2$

$V_2 = 46.9 mL\ HCl$

$203.1 mL\ H_2O$

$250 mL$

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9.11	4.89	7.8×10^{-10}	1.3×10^{-5}	B
7.00	7.00	1.0×10^{-7}	1.0×10^{-7}	N
5.20	8.80	6.3×10^{-6}	1.6×10^{-9}	A

1. H_2O
2. H^+

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